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Chapter 1 Practice 1-5 Guided Problem Solving 1-1

GUIDED PRACTICE PROBLEM 2 (page 324) 2. Sulfur burns in oxygen to form sulfur dioxide. Write a skeleton equation for this chemical reaction. Include appropriate symbols from table 11.1. Analyze Step 1. Identify the relevant concepts. Write the formula for each reactant and each product. Include the common STP state of each substance. Reactants

SECTION 11.1 DESCRIBING CHEMICAL REACTIONS (pages 321-329)

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Practice Problems In your notebook, solve the following problems. SECTION 6.1 ORGANIZING THE ELEMENTS 1. Which element listed below should have chemical properties similar to fluorine (F)? a. Li b. Si c. Br d. Ne 2. Identify each element as a metal, metalloid, or nonmetal. a. fluorine b. germanium c. zinc d. phosphorus e. lithium 3.

Prentice Hall Chemistry Worksheets

guided practice problem 12 (page 513) 12. When 50.0 mL of water containing 0.50 mol HCl at 22.5°C is mixed with 50.0 mL of water containing 0.50 mol NaOH at 22.5°C in a calorimeter, the temperature of the solution increased to 26.0°C.

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GUIDED PRACTICE PROBLEM GUIDED PRACTICE PROBLEM 8 (page 167) 8. Use Figure 6.9 and Figure 6.12 to write the electron configurations of these elements. a. carbon b. strontium c. vanadium Analyze a. What is the number of electrons for each element? C \_\_\_\_ Sr \_\_\_\_ V \_\_\_\_ b. What is the highest occupied energy sublevel for each element, according

SECTION 6.1 ORGANIZING THE ELEMENTS (pages 155-160) (page 155)

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