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Constructed

Response Thoroughly  
and completely  
answer each question  
on a separate piece of  
paper. 8. Consider the  
exothermic reaction  
between reactants A  
and B?  $A + B \rightarrow E$   
(fast)  $E + B \rightarrow C + D$   
(slow) a. What is the  
order with respect to  
reactants A and B? 1,  
2 b.

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Multiple Choice  
Questions Answers -  
Ques. A catalyst  
increases the rate of  
reaction because it (a)  
Increases the  
activation energy

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Questions pertaining to kinetics If you're seeing this message, it means we're having trouble loading external resources on our website. If you're behind a web filter, please make sure that the domains

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From the rate  
expression for the  
following reactions,

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determine their order of reaction and the dimensions of the rate constants.

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Your answers are  
highlighted below.

Question 1 In a  
reaction,  $A + B \rightarrow$   
Product, rate is  
doubled when the

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concentration of B is doubled, and rate increases by a factor of 8 when the concentrations of both the reactants ( A and B ) are doubled, rate law for the reaction can be written as  
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Paper 1

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Important Questions  
for CBSE Class 12  
Chemistry □ Kinetics.  
PREVIOUS YEARS □  
QUESTIONS. 2015  
Short Answer Type  
Questions [1] [2  
Marks] ... Answer :  
Question 29: For a  
chemical reaction  $R$   
 $\rightarrow P$ , the variation in  
the concentration  
 $\ln[R]$  vs. time  $t$  plot is  
given as For this

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reaction (i) What is the  
order of the reaction?

...

Important Questions  
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Chemistry - Kinetics

...

Chemical Kinetics  
Answers: (a)  $8.4 \times 10$   
 $^{-7}$  M/s, (b)  $2.1 \times 10^{-7}$   
M/s SAMPLE  
EXERCISE 14.3  
continued The



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decomposition of  $\text{N}_2\text{O}_5$  proceeds according to the following equation: If the rate of decomposition of  $\text{N}_2\text{O}_5$  at a particular instant in a reaction vessel is  $4.2 \times 10^{-7}$  M/s, what is the rate of appearance of (a)  $\text{NO}_2$ , (b)  $\text{O}_2$ ?

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Kinetics - Free

Response Sample  
Questions 2005 B  
Answer the following  
questions related to  
the kinetics of  
chemical reactions. I<sup>-</sup>  
(aq) + ClO<sup>-</sup> (aq) → H<sup>+</sup>  
+ IO<sup>-</sup> (aq) + Cl<sup>-</sup> (aq)  
Iodide ion, I<sup>-</sup>, is  
oxidized to hypoiodite  
ion, IO<sup>-</sup>, by  
hypochlorite, ClO<sup>-</sup>, in

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basic solution  
according to the  
equation above.

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Response Sample  
Questions  
Chemical Kinetics  
Mastery of  
Fundamentals  
Answers CH353 □  
Prof. Wu Here are  
some questions to  
test your mastery of

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the fundamentals of  
chemical kinetics.

Once you've  
mastered the  
material, you should  
be able to answer  
these questions  
without reference to  
your notes or  
textbook. For  
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Types/Groups of  
problems: Rates of  
Change in Chemical  
Reactions p1 First  
Order Rate Law  
Calculations P9 The  
look of

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concentration/time  
graphs p2 Reaction  
Energy Diagrams,  
Activation Energy,  
Transition States  
P10 Rates: Average  
Rates, Determination  
of Rates from

Test1 ch15 Kinetics  
Practice Problems  
Photochemistry plays  
an important role in  
everyday life. Some of

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the examples are given below: 1. In vision 2. In photosynthesis 3. In the manufacture of many inorganic compounds.

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Chemical kinetics is a study of measuring the rates of chemical



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reactions. This involves calculations, which may require your mathematics knowledge. It is a good teaser to those of us who think we are super good in Chemistry and Math. ... Questions and Answers 1. In a reaction,  $A + B \rightarrow$  Product, rate is doubled when ...

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Test? Trivia Quiz ...

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week of the course. A  
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papers, together with  
worked answers, is  
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Questions - Faculty of  
Science- The ...  
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NCERT InText  
Questions. Question  
1: For the reaction  $R \rightarrow P$ , the concentration  
of a reactant changes  
from 0.03M to 0.02M  
in 25 minutes.

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Calculate the average rate of reaction using units of time both in minutes and seconds.

... Answer: The rate of most of the chemical reactions increases with increase in temperature. The rate ...

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